Matrix Metalloproteinase 2 Inhibition: Combined Quantum Mechanics and Molecular Mechanics Studies of the Inhibition Mechanism of (4-Phenoxyphenylsulfonyl)methylthiirane and Its Oxirane Analogue

Peng Tao,†‡ Jed F. Fisher,§ Qicun Shi,∥ Thom Vreven,∥§ Shahriar Mobashery,∥§ and H. Bernhard Schlegel*†∥

1Department of Chemistry, Wayne State University, 5101 Cass Avenue, Detroit, Michigan 48202, 2Department of Chemistry and Biochemistry, University of Notre Dame, Notre Dame, Indiana 46556, and 3Gaussian, Inc., 340 Quinnipiac Street, Building 40, Wallingford, Connecticut 06492. °Current address: Program in Bioinformatics and Integrative Biology, University of Massachusetts Medical School, Worcester, MA 01605.

Received July 1, 2009; Revised Manuscript Received September 9, 2009

ABSTRACT: The inhibition mechanism of matrix metalloproteinase 2 (MMP2) by the selective inhibitor (4-phenoxyphenylsulfonyl)methylthiirane (SB-3CT) and its oxirane analogue is investigated computationally. The inhibition mechanism involves C–H deprotonation with concomitant opening of the three-membered heterocycle. SB-3CT was docked into the active site of MMP2, followed by molecular dynamics simulation to prepare the complex for combined quantum mechanics and molecular mechanics (QM/MM) calculations. QM/MM calculations with B3LYP/6-311+G(d,p) for the QM part and the AMBER force field for the MM part were used to examine the reaction of these two inhibitors in the active site of MMP2. The calculations show that the reaction barrier for transformation of SB-3CT is 1.6 kcal/mol lower than its oxirane analogue, and the ring-opening reaction energy of SB-3CT is 8.0 kcal/mol more exothermic than that of its oxirane analogue. Calculations also show that protonation of the ring-opened product by water is thermodynamically much more favorable for the alkoxide obtained from the oxirane than for the thiolate obtained from the thiirane. A six-step partial charge fitting procedure is introduced for the QM/MM calculations to update atomic partial charges of the quantum mechanics region and to ensure consistent electrostatic energies for reactants, transition states, and products.

Matrix metalloproteinases (MMPs)1 are zinc-dependent endopeptidases that regulate functions of the extracellular matrix (ECM). MMPs are involved in many biological processes, such as embryonic development (1–3), tissue remodeling and repair (4–6), neuropathic pain processes (7), cancer (8–11), and other diseases (12–15). The physiological activities of the MMPs are strictly regulated by activation of the zymogen forms of the protein (pro-MMPs) (16) and by inhibition by protein tissue inhibitors of metalloproteinases (TIMPs).

Gelatinases A (MMP2) and B (MMP9) are proteolytic enzymes that digest type IV collagens (17). Uncontrolled activities of these two enzymes are implicated in tumor metastasis and angiogenesis. The structures and catalytic mechanisms of MMP2 and MMP9 have been studied extensively (18–24). Since the unregulated activities of these two enzymes have been implicated in many diseases, they are targets for selective inhibitor design (25–31). SB-3CT, one such inhibitor, selectively inhibits MMP2 with high potency and MMP9 with somewhat lower activity (32).

The key event in the inhibition of MMP2 by SB-3CT is enzyme-catalyzed ring opening of the thiirane, giving a stable zinc–thiolate species (33). The previously proposed mechanism for MMP2 inhibition by SB-3CT involved nucleophilic addition of the carboxylate of the active site glutamate to a thiirane carbon (Scheme 1a). This mechanism is preceded with oxirane inhibitors of carboxypeptidase A, a structurally different but mechanistically related protease (34). However, recent experiments from the Mobashery laboratory indicate a different mechanism for SB-3CT (Scheme 1b) (33). In this new mechanism, the carboxylate of glutamate-404 abstracts a hydrogen from the sulfone adjacent to the sulfone (Scheme 1b) (33). This latter mechanism is supported by the observation of a primary deuterium kinetic isotope effect for the methylene group on model systems in solution revealed that the barriers of ring opening initiated by deprotonation are lower than those of the glutamate addition mechanism (35). These calculations also reproduce the observed primary kinetic isotope effect. Given these facts, the mechanism involving covalent modification of the glutamate was not pursued in current study.

In preparation for the present investigation, we studied the conformational preferences and stereochemical aspects of the
Scheme 1: MMP2 Inhibition Mechanisms by SB-3CT: (a) Previously Proposed Mechanism; (b) Current Mechanism

Scheme 2: Structures of SB-3CT (3) and Its Analogues (1, 2, and 4)

Coupled deprotonation and ring opening in solution of 2-(methylsulfonylmethyl)thiirane and -oxirane (1 and 2 in Scheme 2) as models for SB-3CT (3) and its oxirane analogue (4). Proton abstraction can occur either syn or anti to the three-membered ring, and the sulfone group exerts a comparable stereoelectronic effect in the thiirane and in the oxirane. Since no crystal structure is available for the MMP2-SB-3CT complex, the structure and stability of the complex were assessed by docking of SB-3CT into the MMP2 active site, followed by molecular dynamics studies. Then the details of the deprotonation/ring-opening mechanism for inhibition were examined by combined quantum mechanics and molecular mechanics (QM/MM) methods.

**COMPUTATIONAL METHODS**

**Docking and Molecular Dynamics Studies of the MMP2-SB-3CT Complex.** Since the structure of the non-covalent MMP2-SB-3CT complex is not experimentally accessible, SB-3CT was docked in the active site of the crystal structure for the Ala404 mutant of MMP2 (PDB code 1CK7) (18). Ala404 was computationally mutated to Glu404, the catalytic base in the MMP2 active site. The propeptide domain (residues 31–115) was deleted, as would be the case in the active form of MMP2. The resulting MMP2 enzyme includes residues 116–449, two zinc metal ions (Zn990 and Zn991), and three calcium ions. SYBYL (TRIPOS 7.3) (36) was used to prepare the structures of the inhibitors. DOCK (version 5.4; UCSF) (37) was employed to dock the inhibitor in the active site, using electrostatic and van der Waals forces to score the acceptor–inhibitor interactions. The docked MMP2-SB-3CT complex was immersed in a water solvent box through energy minimization and thermodynamic equilibration procedures (using XLEAP from AMBER 9). AMBER force field (parm99) was used to describe the whole system, including zinc ions. The force field parameters for zinc (38) are listed in the Supporting Information. During these stages, position constraints were enforced for the atoms in the three histidine residues surrounding each of the zinc cations, Glu404, and SB-3CT substrate with harmonic potentials of approximately 1 Å width and force constants of 50 kcal·mol⁻¹·Å⁻². Furthermore, a distance constraint was added between Zn990 and the nitrogen atom of each histidine at the value given in the crystal structure, using a harmonic potential of width 0.2 Å and force constants of 1000 kcal·mol⁻¹·Å⁻². A total of 2.0 ns of molecular dynamics (MD) simulation was carried out. Snapshots were extracted every 0.5 ps. The conformation of the complex was analyzed for each of the 4000 snapshots.

**QM/MM Studies of the MMP2-SB-3CT Complex.** The initial structure for the QM/MM calculations of the reactant complex was prepared using the AMBER software suite, version 9 (39). Since experiments show that (R)-SB-3CT is slightly more active than (S)-SB-3CT in terms of $k_{on}$ (40), the R stereoisomer was chosen for the QM/MM studies. The selected MMP2-(R)-SB-3CT complex from the docking and MD simulations (approximately 66000 atoms) was subjected to geometry optimization using the SANDER program from AMBER 9 (approximately 12000 conjugate gradient steps). During this optimization, a constraint was applied to the distance between an oxygen atom of the carboxylate group of Glu404 and a hydrogen atom of the CH₂ group adjacent to the thiirane ring of SB-3CT to ensure that the optimized geometry would be appropriate for the abstraction step. The AMBER optimized geometry was used as a starting point for the QM/MM calculation. The active site Zn²⁺, His403, His407, His413, Glu404, and the SB-3CT inhibitor were defined as core residues. The water molecules that are either within 3 Å of protein or within 12 Å of the core residues were kept for the QM/MM calculations. All other water molecules were deleted to facilitate the computation. The final system subjected to QM/MM calculations comprised approximately 8800 atoms. In the QM/MM calculations, all residues within 6 Å of the core residues were allowed to move without any constraints, while all other residues were frozen. The oxirane optimizations started from the thiirane optimized geometries. Some care is needed to ensure that the hydrogen-bonding pattern in the solvent water molecules is the same for the reactant complex and the transition state so that minor changes in the solvent do not overwhelm differences in the barrier heights.

A two-layer ONIOM method (41–50) was used for the QM/MM study of the inhibition mechanism of SB-3CT and its oxirane analogue. The zinc ion, the three imidazole rings from His403, His407 and His413, the CH₂CO₂⁻ part of the Glu404 side chain, the thiirane or oxirane with the SO₂CH₂ group, and two water molecules coordinating with zinc in the MD simulation were included in the QM region (49 atoms). The QM part of the system was described at the B3LYP/6-31G(d) level of density functional theory. Comparison with benchmark calculations at the CBS-QB3 level for the deprotonation/ring-opening reactions in the gas phase showed that B3LYP performed significantly better than HF or MP2 calculations (35). The MM part of the system was described using the parm96 parameters of the AMBER force field (51). A mechanical embedding scheme was used for geometry optimization (electrostatic interactions between the QM and MM regions are handled by MM in this approach). Transition state searches used the quadratically coupled QM/MM geometry optimizer (48) implemented in the development version of the GAUSSIAN package (32). The optimizer explicitly calculates the transition vector, which
through the quadratic coupling between the regions is not restricted to the QM region of the system. The molecular mechanics contributions to the second derivative matrix are always evaluated analytically, which makes the optimization procedure quite reliable. This method has been applied previously to calculate transition states in other enzymatic systems (53–55). The final QM/MM energies reported are based on electronic embedding (49) single point calculations at the ONIOM(B3LYP/6-311+G(d,p):AMBER) level of theory using the ONIOM(B3LYP/6-31G(d):AMBER) optimized geometries. All ONIOM calculations were carried out with the development version of GAUSSIAN (52).

The RESP (restrained electrostatic potential) program (56, 57) was used to fit partial charges to the electrostatic potential (ESP) grid generated by gas phase calculation. In this study, Glu404 accepts a proton from the inhibitor. This changes the partial charge distribution of substrate and Glu404 significantly. Therefore, using fixed partial charges for the substrate and protein along the reaction path could reduce the accuracy of the ONIOM energies. To address this issue, we developed a six-step procedure for fitting a consistent set of partial charges for the reactant, TS, and product: (1) A preliminary set of partial charges was obtained for the substrate in the gas phase using the RESP procedure. (2) The reactant, TS, and product were optimized by ONIOM calculations using mechanical embedding with the partial charges. (3) The QM atoms of the reactant, TS, and product from the ONIOM geometry optimizations in the second step were used to obtain an improved set of partial charges using the RESP procedure, and the hydrogen atoms added to cap the dangling bonds are constrained to have zero charges. (4) The reactant, TS, and product were optimized in the active site by ONIOM calculations using mechanical embedding with the improved charges. (5) Steps 3 and 4 are repeated until convergence is achieved. The convergence criterion is that the total ONIOM energy difference between the last two round optimizations is less than 0.1 kcal/mol. (6) The converged geometries and charges from the last step are used for single point calculations with electronic embedding. This typically required four to six rounds of optimization to obtain partial charges appropriate for the reactants, TSs, and products in their optimal geometries in the active site. With this method, the total charge on the QM atoms is conserved along the reaction path, and the changes in the electrostatic interaction energy in going from reactants to TSs to products are reproduced properly. This procedure was followed in the present work since the charge distribution was suspected to be critical for reliable results. For other studies, the dependence of the results may not be as sensitive to the charges, and iterating may not be necessary.

RESULTS AND DISCUSSION

Docking and Molecular Dynamics Studies. Both the (R)- and (S)-SB-3CT enantiomers are effective inhibitors of this enzyme (40). Since the (S)-SB-3CT enantiomer is the slightly less reactive, (R)-SB-3CT was chosen for the initial computational study. (R)-SB-3CT was docked into the active site before being subjected to MD simulation. The MMP2·SB-3CT complex remains stable during the 2 ns MD simulation. The key electrostatic interactions that stabilize SB-3CT in the MMP2 active site involve the attraction between zinc and the sulfur of the thirane, a hydrogen bond between one oxygen of the sulfone and the amide hydrogen of Leu191, and hydrophobic interactions between the phenoxyphenyl ring and the residues of the S1′ pocket (58–60).

The MD trajectory of MMP2·(R)-SB-3CT was analyzed to select a representative conformer for the QM/MM calculations. The most populated ranges are 2.5±0.1 Å for distance between Zn and thirane S (d1) and 1.9±0.1 Å for distance between Leu191 amide H and one oxygen of the sulfone group (d2). The conformer with both d1 and d2 within the most populated ranges and the shortest distance between the Glu404 oxygen and the pro-S hydroxyl of (R)-SB-3CT (d3) was chosen as a starting point for QM/MM calculations. The chosen conformer has d1 = 2.53 Å, d2 = 1.84 Å, and d3 = 2.26 Å (Figure 1).

(S)-SB-3CT was also subjected to MD simulation as comparison. An initial structure for the S enantiomer bound in the active site of MMP2 was generated from the MD-optimized R enantiomer structure by inversion of configuration and subjected to 2 ns MD simulation. The same constraints as used for the R enantiomer were employed in the MD simulation of the S enantiomer. The resulting conformations maintain the coordination between the zinc cation and the thirane ring sulfur (d1 = 2.40 Å) and between a sulfone oxygen and the Leu191 amide hydrogen (d2 = 1.80 Å), as observed in the R enantiomer. A conformation having a distance of 2.36 Å (d3) between the Glu404 and the pro-R hydrogen of the methylene carbon was extracted (Supporting Information Figure S1). Of the distance between the oxygen and hydrogen atoms, the molecular dynamics samplings showed similar distributions over the 2 ns MD simulation for both the R and S enantiomers but did indicate diastereomeric preference of the glutamate for the pro-S hydrogen in the (R) enantiomer and for the pro-R hydrogen in the S enantiomer. The differences in the conformations of the R and S enantiomers are primarily (1) rotation of the thirane ring along the C–C bond and (2) rotation of the phenoxyphenyl rings along the adjacent CS bond (Figure 1; the S enantiomer is shown in Supporting Information Figure S1).

FIGURE 1: Structure of the MMP2·(R)-SB-3CT complex from the MD simulation. The distance between zinc and thirane sulfur is 2.53 Å, which is close to the most populated value in the MD simulation. Residues are shown in ball-and-stick representation with atom colored according to atom types (H, C, N, O, S, and Zn shown in white, cyan, blue, red, yellow, and gray, respectively). The same color scheme is used in all of the figures.
QM/MM Studies. The docking study and the subsequent MD samplings generated a reliable structure of the MMP2 and SB-3CT complex. Conformations with a short distance between the hydrogen of the α-methylene and Glu404 were sampled for the initiation of the deprotonation reaction. However, the ring-opening reaction coupled with deprotonation cannot be studied by the molecular mechanics level of theory used in docking and MD. Therefore, QM/MM methods were used to carry out further calculations.

(a) Water or Hydroxide in the Active Site. The MD study of MMP2-SB-3CT shows two water molecules coordinated with the zinc in the active site (Figure 1). One water molecule is located between the zinc and Glu404 (Wat1), and the second is in contact with bulk solvent (Wat2). The QM/MM calculations show spontaneous proton transfer from Wat1 to the Glu404 carboxylate, giving a hydroxide anion coordinated to the zinc (Figure 2; geometry parameters are given in Supporting Information Figure S2). A similar proton migration from a thiol group to glutamate, to produce thiolate as a zinc ligand in pro-MMP9, was observed in the previous QM/MM study on the activation of pro-MMP9 (16). The zinc is coordinated with hydroxide and the three histidines. Neither SB-3CT nor the carboxylate side chain of Glu404 is coordinated with the zinc in this structure (5.13 and 5.08 Å). It is noteworthy that in the same study of MMP9 (16) a water molecule enters the active site, inserting itself between Glu402 and zinc during the activation process of pro-MMP9. Zinc has tetrahedral coordination with a water molecule and three histidines, and the water forms a strong hydrogen bond with the glutamate.

Since Glu404 needs to be in a deprotonated state to initiate the reaction with the inhibitor, additional QM/MM calculations were carried out with hydroxide bound to the zinc and Glu404...
deprotonated (Figure 3 and Supporting Information Figure S3). The proton abstraction/ring-opening energies are quite endothermic: 15.8 and 6.6 kcal/mol for the thiirane 3 and oxirane 4, respectively. The hydroxide coordinates with zinc in both the reactant and product structures. Glu404 forms a hydrogen bond with the hydroxide and coordinates with zinc in the reactants. However, Glu404 moves away from zinc in the products. Neither the thirane nor oxirane coordinates with zinc in the reactants. The overall reactions are endothermic because the ring-opening products do not coordinate with the zinc.

These QM/MM calculations show that if a water molecule or hydroxide anion is in the active site shielding zinc from Glu404, the inhibitor cannot interact with the zinc, and the pathways are endothermic and not thermodynamically feasible. Therefore, subsequent QM/MM calculations were carried out without a water molecule between zinc and Glu404.

(b) Deprotonation and Ring Opening of the Inhibitor. In the initial structure for further QM/MM calculations, Glu404 and three histidines coordinate with zinc (see Figure 4). The water molecule (Wat2) is not coordinated to the zinc but is open to the solvent. The orientation and key intermolecular contacts between (R)-SB-3CT and the MMP2 active site closely resemble the crystallographic structures seen for other MMP inhibitors (58). Particularly, the phenoxyphenyl side chain is located as expected in the S1' pocket, and the customary strong (1.9 Å) hydrogen bond from the backbone NH of Leu191 to the pro-S oxygen of the sulfone is preserved (Figure 4). As is also expected, the second oxygen of the sulfone is solvent-exposed. In the QM/MM optimized reactant complex of SB-3CT, the zinc is coordinated with the three histidines, the Glu404 carboxylate, and the thirane sulfur. No crystal structure is available for SB-3CT bound to a matrix metalloproteinase; however, this coordination agrees with the modeled complex structure between MMP9 and SB-3CT (58). These observations support the absence of a water molecule between Glu404 and the zinc ion. Of particular interest to the deprotonation mechanism is the conformation of the bound SB-3CT. In addition to the intermolecular interactions enumerated above, a stereoelectronic effect governs the orientation of the phenylsulfone segment, wherein the π-orbital of the

Figure 4: MMP2-(R)-SB-3CT complex structure from QM/MM calculation (the same geometry as 3-R in Figure 5).
The reactant, transition state, and product structures for the inhibition of MMP2 by (R)-SB-3CT are shown in the top row of Figure 5, and selected geometrical parameters are given in Supporting Information Figure S4. Those for the oxirane analogue are shown in the bottom row of these figures. In the reactant complex of SB-3CT (3-R), an oxygen from the carboxylate group of Glu404 and the sulfur from the thirane coordinate with the zinc at 1.99 and 2.91 Å, respectively, while the water is not coordinated (3.58 Å). For the oxirane reactant, 4-R, the interaction with the active site is very similar to the thirane reactant. The QM/MM energies with electronic embedding are given in Table 1 and plotted in Figure 6.

The QM/MM calculations of the thirane transition state, 3-TS in Figure 5 and Supporting Information Figure S4, the transferring proton is 1.14 Å from the acceptor oxygen and 1.57 Å from the donor carbon. The breaking C–S bond of the thirane is elongated to 2.03 Å in 3-TS. The Glu404 side chain has moved away from zinc in order to abstract the proton. One oxygen of the sulfone accepts a hydrogen bond from the backbone NH of Leu191, as shown in Figure 4. In the TS of the oxirane analogue (4-TS), the proton transfer is earlier along the path, and the ring opening is similar to the thirane system. The thirane sulfur and oxirane oxygen are strongly coordinated to the zinc, but the glutamate is not.

Although the bound conformation of (R)-SB-3CT in the reactant complex shows a near-staggered conformation with respect to the mechanistically critical C–C dihedral angle for the ring-opening reaction, the spatial placement of Glu404 implicates a syn relationship for proton abstraction and concomitant thirane opening. In a separate study, the model molecule of SB-3CT (1) reacting with acetate as the Brønsted base in solution, the C–S bond dihedral in the syn-elimination TS is −34° (35). In the MMP2 QM/MM TS (Figure 5 and Supporting Information Figure S4, 3-TS), this dihedral angle is 45°. To estimate the energetic cost imposed by this dihedral angle difference, the syn-elimination reactant complex and TS of 1 were recomputed by constraining the syn-elimination dihedral angle to the 45° seen in QM/MM structures. The two transition states computed in methanol solution (syn dihedral of −34° and 45°) are compared in Figure 7. The enthalpy of TS with dihedral angle constraint is 4.1 kcal/mol higher than the optimal syn-elimination TS in methanol.

In the products, 3-P1 and 4-P1, the thiolate sulfur and the alkoxide oxygen remain tightly coordinated to the zinc. The protonated Glu404 stays away from the zinc. As in the transition states, the zinc has tetrahedral coordination, and the water molecule does not interact with the zinc. Comparing the structures of the reactant complex and TS of SB-3CT (3) in the MMP2 active site (Figure 5 and Supporting Information Figure S4, 3-R and 3-TS, respectively), it is apparent that Glu404 moves a considerable distance from the zinc toward the inhibitor from the reactant complex to TS. Clearly, the carboxylate of Glu404 must disengage from the zinc. The energy surface between the TS and the reactant and product complexes of MMP2–(R)-SB-3CT and its oxirane analogue was examined by starting near the TS and minimizing the energy in the reactant and product directions. The energies along these optimization paths confirm that no additional barriers are encountered between the TS and the reactant or the product complexes for both 3 and 4 (see Figures S5 and S6 in the Supporting Information).

The barriers for the thirane and oxirane ring opening at the active site of MMP2 are 19.9 and 21.5 kcal/mol, respectively, and the corresponding reaction energies are −21.1 and −13.1 kcal/mol. These values indicate a role for the enzyme in this base-mediated elimination reaction in stabilization of the ring-opening product, therefore yielding much more favorable reaction energies.
for the inhibitors in the active site compared to the model systems in solution (35). The results show that the reactions of both 3 and 4 within the MMP2 active site are kinetically feasible and thermodynamically favorable.

The deprotonation-initiated ring-opening TSs identified for 3 and 4 at MMP2 active site closely resemble the TSs of model systems in solution (35). The calculated kinetic isotope effect for the thirane model system (kD/H = 5.0) is in good agreement with experimental results (kD/H = 5.0) (35). The evaluation of a kinetic isotope effect in a QM/MM calculation would necessitate the computation of a full set of vibrational frequencies or extensive sampling of the potential energy surface for the reactant and the transition state. Such calculations are not practical at this time but may become feasible in the future.

When complexed to MMP2, the reaction of thirane 3 is calculated to be about 8 kcal/mol more exothermic than the reaction of oxirane 4. The difference in exothermicity of 3 and 4 shows that the ring-opening reaction of SB-3CT is thermodynamically more favorable than its oxirane analogue. In the present computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

Comparison with the calculations in solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The combination of a lower inhibitor of MMP2 that binds poorly to the active site oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The difference in the barrier may be larger than 1.6 kcal/mol. The comparison with the calculations in solution suggests that the reaction of SB-3CT is thermodynamically more favorable than its oxirane analogue. In the present computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.


The deprotonation-initiated ring-opening TSs identified for 3 and 4 at MMP2 active site closely resemble the TSs of model systems in solution (35). The calculated kinetic isotope effect for the thirane model system (kD/H = 5.0) is in good agreement with experimental results (kD/H = 5.0) (35). The evaluation of a kinetic isotope effect in a QM/MM calculation would necessitate the computation of a full set of vibrational frequencies or extensive sampling of the potential energy surface for the reactant and the transition state. Such calculations are not practical at this time but may become feasible in the future.

When complexed to MMP2, the reaction of thirane 3 is calculated to be about 8 kcal/mol more exothermic than the reaction of oxirane 4. The difference in exothermicity of 3 and 4 shows that the ring-opening reaction of SB-3CT is thermodynamically more favorable than its oxirane analogue. In the present computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The deprotonation-initiated ring-opening TSs identified for 3 and 4 at MMP2 active site closely resemble the TSs of model systems in solution (35). The calculated kinetic isotope effect for the thirane model system (kD/H = 5.0) is in good agreement with experimental results (kD/H = 5.0) (35). The evaluation of a kinetic isotope effect in a QM/MM calculation would necessitate the computation of a full set of vibrational frequencies or extensive sampling of the potential energy surface for the reactant and the transition state. Such calculations are not practical at this time but may become feasible in the future.

When complexed to MMP2, the reaction of thirane 3 is calculated to be about 8 kcal/mol more exothermic than the reaction of oxirane 4. The difference in exothermicity of 3 and 4 shows that the ring-opening reaction of SB-3CT is thermodynamically more favorable than its oxirane analogue. In the present computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.

The computational model, the ring-opening barrier of 4 is higher than 3 by 1.6 kcal/mol. This would indicate that 3 is only a factor of 15 more reactive than is 4. However, calculations in aqueous solution using a polarizable continuum model yield 6 kcal/mol for the difference in the barrier heights (35). One of the limitations of QM/MM calculations at the present state of the art is the inability to carry out extensive sampling. Since the active site is open to the solvent, fluctuations in hydrogen bonding to the solvent could add an uncertainty of ±2 kcal/mol or more to the difference in barrier heights obtained by local optimization. Comparison with the calculations in solution suggests that the difference in the barrier may be larger than 1.6 kcal/mol. The oxirane analogue is known to be a weak, linearly competitive inhibitor of MMP2 that binds poorly to the active site.


25. SYBYL 7.3, Tripos Inc., 1699 South Hanley Rd., St. Louis, MO 63144.


